

13.2 - properties of liquids

Title: Properties of Liquids	Surface Tension
	Capillary Action
	Evaporation or Vaporization
	Vapor Pressure

→ form of evaporation:
sublimation
solid → vapor

→ volatile - substances that evaporate readily (↑ vapor pressure @ room temp)

→ condensation - Vapor (strikes surface +) returns to liquid

→ equilibrium equation =
liquid $\xrightleftharpoons[\text{condensation}]{\text{evaporation}}$ Vapor

13.3 - b.p. + m.p. (see study guide)

13.2 - properties of liquids

↑ attractive forces b/w molecules = ↑ surface tension

- the resistance of a liquid to ↑ its surface area
- when water not falling under gravity, forms sphere to minimize the ratio of surface area to volume

if forces b/w liquid + container > those w/in liquid itself
= liquid will climb

- spontaneous rising of a liquid in a narrow tube
- results from the cohesive forces w/in liquid + the adhesive forces b/w the liquid + the walls of the container

endothermic physical process
(when leave, liquid becomes cooler)

- the escape of molecules from the liquid state to the gas or vapor state
- molecules that have greater-than-avg. KE can overcome attractive forces + break away

→ form of evaporation: sublimation
solid → vapor

→ volatile - substances that evaporate readily (↑ vapor pressure @ room temp)

Vapor pressure of liquid is independent of amt. of liquid + vapor present, but ↑ as temp. ↑

- @ equilibrium, molecules in the vapor exert pressure, this pressure with its liquid is known as vapor pressure of the liquid
- measure of "escaping" tendency

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