

L97 - What's in a Mole? - ACTIVITY

Part 1: Molar Mass

Chemical Formula	Molar Mass (g/mol)	Mole of what?	Equivalent to:
Cu(s)	63.54	Cu atoms	50 ft of 20-gauge copper wire
O ₂ (g)	32.00	O ₂ molecules	_____ L oxygen gas at STP
Ni(s)		Ni atoms	12 nickel coins
Al(s)		Al atoms	2 aluminum cans
H ₂ O(l)		H ₂ O molecules	_____ water
He(g)			_____ oxygen gas at STP
NaCl(s)			
Hg(l)			_____ mercury
Fe(s)			
C ₁₂ H ₂₂ O ₁₁ (s)			

Please complete
all parts



Part 2: Mole Challenge

Station #	Total Measured Mass (e.g., vial + cap + sample) (g)	Measured Mass of Vial and Cap (g)	Calculated Molar Mass (g/mol)	Physical Description	Element Symbol	Molar Mass from Periodic Table (g/mol)
1						
2						
3						
4						
5						
6						
7						
8						
9						
10						
11						
12						