

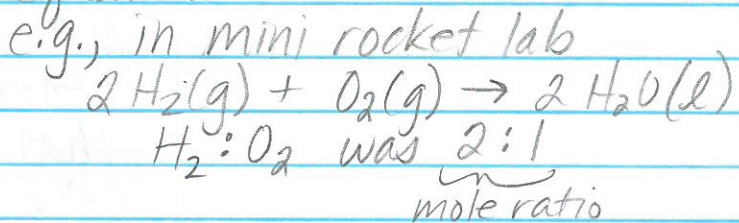
L91 Mole Ratios

TODAY'S
DATE

E.Q.: How can you convert all the reactants to products?

mole ratio → the ratio of the moles of one reactant or product to the moles of another reactant or product in a balanced chem equation

e.g., in mini rocket lab



$\text{H}_2 : \text{O}_2$ was $\underbrace{2:1}$
mole ratio

BIG IDEA: To get the max amt. of product from a rxn, reactants must be mixed in correct proportions

If you mix the reactants in a ratio other than the mole ratio specified by the balanced chem eq., a rxn will still occur, but one of the reactants will run out and the other will have some left over

limiting reactant (or limiting reagent) - the reactant that runs out first in a chem rxn - it is the reactant that limits the amt. of product that can be produced in the rxn